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### ► To cite this version:

Rabea Bouayad, Jérôme Bikard, Jean-François Agassant. Injection of thermoset foam: comparison between simulation and experiment. 11th ESAFORM Conference on Material Forming, Apr 2008, Lyon, France. pp.683-686, 10.1007/s12289-008-0307-6 . hal-00510258

**HAL Id: hal-00510258**

**<https://hal-mines-paristech.archives-ouvertes.fr/hal-00510258>**

Submitted on 18 Aug 2010

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# Injection of thermoset foam: comparison between simulation and experiment

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**ABSTRACT:** The quality (cellular homogeneity, mechanical properties) of polyurethane foam's structures mainly depends on the manufacturing process, during which two concomitant (principal) exothermic chemical reactions take place: the first one creates CO<sub>2</sub> into the fluid matrix (germination of bubbles, expansion and coarsening of the foam) and the second one leads to the polymerization. In order to validate a model developed at CEMEF, an original experiment (RheoFoam System) has been created. It consists in an instrumented injection mould (closed or opened cylindrical cavity) in which the viscoelastic foam inflates. It allows measuring simultaneously the evolution of some technological parameters (the rise of the foam, the pressure distribution on the bottom of the mould and the temperature evolution inside the foam) which are a macroscopic signature of the evolution of the cellular microstructure. These temperature and pressure fields are then compared to those obtained using the numerical simulation. The results are discussed.

**Key words:** foam expansion, chemical reactions, diphasic medium, finite element modeling, experiments

## 1 INTRODUCTION

Flexible polyurethane foams are usually applied in car seat parts. Optimisation of the manufacturing process, as well as part quality may be improved through numerical modelling. Flexible PU foams are produced in a one shot process in which (poly)isocyanate, polyols and water are mixed simultaneously with suitable stabilisers, catalysts and cell-size control agents. The chemical reactions begin immediately, with foam rise starting a few seconds after mixing and being completed in a matter of minutes. Curing continues for several hours, eventually leading to a solid cellular material [1,2].

The two primary reactions are the curing reaction, which leads to formation of polyurethane, and the expansion reaction, producing polyurea and carbon dioxide, with simultaneous expansion of CO<sub>2</sub> bubbles (foaming) and polymerization of the mixture. The first step of the expansion is bubble nucleation, where CO<sub>2</sub> molecules dissolved in the mixture initiate micro-bubbles, under the effect of pressure decrease at the exit of the injection syringe into the mould. At a macroscopic scale, the nuclei can be modelled by an initial porosity of the gas/polymer mixture. Two mechanisms have to be distinguished at this scale: the expansion by difference of pressure and by gas creation. During

fabrication of this foam, the rheological properties of its skeleton evolve from a viscous liquid to a viscoelastic (or elastic) solid.

The objective of the present work is the identification of several parameters of a numerical model of foam expansion developed at CEMEF [3] by comparison between experiments in a cylindrical mould and numerical computations.

Section 2 presents the physical assumptions and the equations of the model. The first part of section 3 is devoted to the presentation of the experiment used to identify some parameters of the model. The last part of section 3 shows a comparison between these experiments and numerical simulations.

## 2 MODELLING

The objective of the model is to predict at a macroscopic level the expansion of the foam (corresponding to domain  $\Omega_m$  characteristic of the foam sample shown on Fig. 1.) into a mould (domain  $\Omega$ ). It is based on the conservation equations (mass and stress balance), which are written by considering that  $\Omega_m$  is a homogenized medium (polymer  $\Omega_l$  + gas bubbles  $\Omega_{bi}$ ,  $i=1,2,\dots,N_{\text{bubbles}}$ ). The interactions between polymer and gas bubbles are described by the evolution of the porosity  $\phi$  [1,4]. The free surface between  $\Omega_m$  and air  $\Omega_a$  (see Fig. 1.) is a result of the numerical model

(section 3.2).

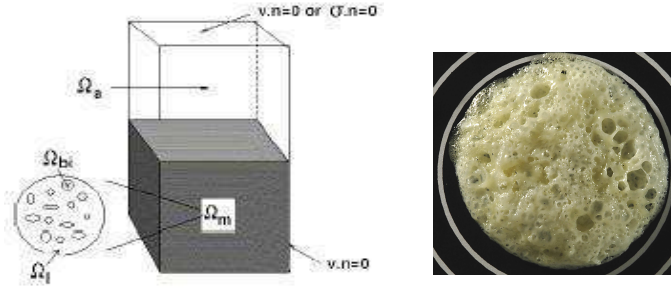


Fig. 1. left: Scheme of the mould containing PU mixture (gas+polymer,  $\Omega_m$ ) and the air ( $\Omega_a$ ). The contact with the mould is assumed perfectly sticking ( $v \cdot n = 0$ ) ; right: a sample of PU foam (average radius 2cm).

Due to the fact the chemical reactions are strongly exothermal, the model takes into account the thermo-mechanical couplings.

### 2.1 Global mass conservation

The global mass conservation leads to the following local equation in  $\Omega_m$ :

$$\nabla \cdot v = \frac{1}{1-\phi} \frac{d\phi}{dt} \quad (1)$$

where  $v$  is the expansion velocity of the foam.

### 2.2 Kinetics evolution laws

Gas creation and curing reactions are governed by chemical kinetics, whose conversion rates are supposed to follow evolutions law [1,2]. Concerning gas creation, the following Kamal law is assumed [5] in  $\Omega_m$ :

$$\frac{\partial \alpha}{\partial t} + v \cdot \nabla \alpha = \lambda_g(T) \alpha^{\mu_g} (1-\alpha)^{v_g} + \nabla \cdot (D_g(T) \nabla \alpha) \quad (2)$$

where  $\alpha$  is the characteristic rate of gas creation,  $\lambda_g^{-1}$  its characteristic time (depending on the temperature) and  $\mu_g$  and  $v_g$  the exponents of this reaction.  $D_g$  represents the diffusion coefficient of gas into the polymer. Assuming a perfect gas law in the bubbles, the porosity development can be macroscopically written [3] in  $\Omega_m$  by:

$$\frac{d\phi}{dt} = \phi(1-\phi) \left[ \frac{1}{\alpha} \frac{d\alpha}{dt} - \frac{1}{p} \frac{dp}{dt} + \frac{1}{T} \frac{dT}{dt} \right] \quad (3)$$

where  $p$  is the hydrostatic pressure in the foam. The polymerization reaction leads to the viscosity increase of the matrix as a function of the

temperature up to a gel point. The curing rate is supposed to follow also a Kamal law [5] in  $\Omega_m$ :

$$\frac{\partial \beta}{\partial t} + v \cdot \nabla \beta = \lambda_p(T) \beta^{\mu_p} (1-\beta)^{v_p} \quad (4)$$

$\beta$  is the characteristic rate of the cure,  $\lambda_p^{-1}$  its characteristic time and  $\mu_p$  and  $v_p$  two exponents of the reaction.

### 2.3 Quasi-static balance equations

Experimentally, the global expansion of PU foams is slow (the strain rate is about  $10^{-2} \text{ s}^{-1}$ ). Assuming a quasi-static evolution, the balance equations reduced to [3]:

$$\begin{cases} \text{div } \sigma = 0 & \text{in } \Omega \\ \sigma = \sigma_l = 2\eta(\dot{\gamma}, \phi, \beta) \cdot \left[ \bar{\nabla}_{Sym} \bar{v} - \frac{1}{3} I : \bar{\nabla}_{Sym} \bar{v} \right] - pI & \text{in } \Omega_m \\ \sigma = \sigma_a = -p_a I & \text{in } \Omega_a \\ \sigma \cdot n = 0 & \text{on } \Omega_a \cap \Omega_m \end{cases} \quad (5)$$

where  $\sigma$  is the Cauchy stress tensor,  $\varepsilon(v)$  the strain rate tensor,  $\dot{\gamma} = \sqrt{2\varepsilon(v) : \varepsilon(v)}$  the second invariant of the strain rate tensor,  $\eta$  the viscosity of the mixture,  $p_a$  the pressure in the air. The interface conditions assume the continuity of the normal velocity and the normal stress (the interfacial tension is neglected, this is a strong hypothesis).

### 2.3 Energy balance

From thermodynamical considerations, the heat equation can be written by:

$$\rho C(\alpha) \frac{dT}{dt} = \nabla \cdot (\lambda_T(\alpha) \nabla T) + \delta H_\alpha \frac{d\alpha}{dt} + \delta H_\beta \frac{d\beta}{dt} + \eta \dot{\gamma}^2 + T \frac{\partial \sigma}{\partial T} : (\nabla_S \bar{v})_{dev} - p \nabla \cdot \bar{v} \quad (6)$$

where  $\delta H_{\alpha, \beta}$  are the enthalpies of both reactions,  $\rho C$  is the heat capacity and  $\lambda_T$  the thermal conductivity.

### 2.4 Rheological coupling

The matrix is considered as a shear-thinning fluid, whose behaviour is expressed by a Carreau law:

$$\eta(\dot{\gamma}, \phi, \beta) = \eta_0(T) \left( 1 + a(T_{ref})^2 \dot{\gamma}^2 \right)^{\frac{m-1}{2}} f(\phi) g(\beta) \quad (7)$$

where  $\eta_0$  is the Newtonian plateau viscosity,

following a classical Arrhénius law,  $a$  is a characteristic time and  $m$  the power-law exponent. Expansion and curing reactions will modify the viscosity of the fluid through two functions  $f$  and  $g$ , which follow the model developed by Castro and Macosko [6]:

$$f(\phi) = f_0 - f_1\phi + f_2\phi^2 \quad \text{and} \quad g(\beta) = \left( \frac{\beta_{gel} - \beta}{\beta_{gel}} \right)^{-n_g} \quad (8)$$

where  $\beta_{gel}$  is the gel point and  $f_0, f_1, f_2$  and  $n_g$  are positive constants.

### 3 ESTIMATION OF SOME RHEOLOGICAL PARAMETERS

The characteristic value of several parameters can be found in literature (see Table 1). Some of them can be identified using dynamic rheology experiments [7].

parameter	$\lambda_g^{-1}$	$\mu_g$	$\lambda_p^{-1}$	$\mu_p$	$\nu_p$	$\eta_0$	$f_0$
Range	1-2	0-2	1-10	0-2	0-2	$10^2$ - $10^3$	1
	min		min			mPa.s	
parameter	$F_1$	$F_2$	$\beta_{gel}$	$n_g$	$a$	M	
Range	1-10	0-1	0,9-1	1-5	0-10 s	0-1	

Table. 1. Values of parameters used in the model.

Resolution of the model in an axisymmetric configuration and comparison with well instrumented experiments allow to identify more precisely these parameters.

#### 3.1 Experiment

Fig. 2 shows the experimental cylindrical mould: the cylinder is closed after the components of PU have been mixed and put at the bottom of the mould. Pressure and temperature are recorded. Gas outlet during the expansion (gas initially present in the mould + degassed  $\text{CO}_2$ ) is also measured and correlated to the expansion velocity of the foam.

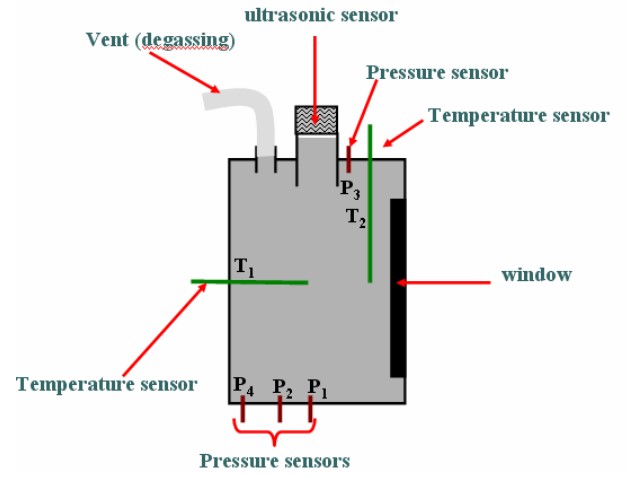
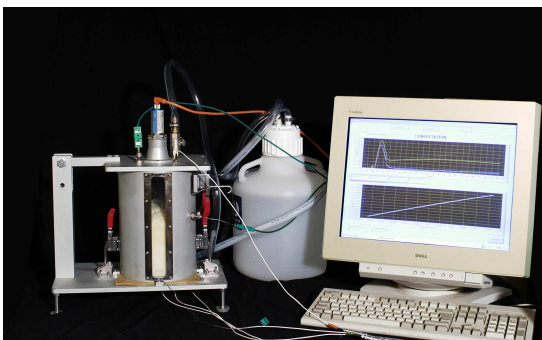


Fig. 2. Photograph and Scheme of the experimental mould.

After opening of the mould, one recovers a typical foam cylinder as shown on Fig. 3.



Fig. 3. Photograph of a typical foam cylinder manufactured using the RheoFoam.

#### 3.2 Numerical resolution and comparisons

Equations (1-8) are highly coupled and non-linear. The numerical method is based on a splitting technique (used by [8] in the case of the microscopic simulation of PU expansion): at one time step, knowing  $T$ ,  $\phi$  and  $\beta$ , velocity and pressure fields are first determined through a mixed finite element method, verifying stability conditions [9]. The velocity is then used to compute temperature  $T$ , gas production  $\alpha$  and solidification rate  $\beta$ . Finally, the moving interface between the gas-liquid mixture and air is computed, introducing a characteristic function of the mixture as additional unknown in each time interval [10] and solved by a volume of fluid method (V.O.F.) associated with a Space-Time Discontinuous Galerkin technique. The model has been implemented in the Rem3D® software [3].

Using characteristic values of parameters (see table 1), a numerical simulation is performed under the same conditions than the experiment described above, and then optimized in order to obtain a computed foam size equivalent to the experimental



one. At that time the optimization loop The expansion results are shown on fig. 4: the gas rate is plotted on a cross section of the numerical cylinder respectively for 80, 100 and 150 s. (see also Figure 5). Due to the fact no nucleation mechanism has been taken into account, the gas rate can be directly correlated to the characteristic length of the foam cells. One observes that the larger ones are located close to the top of the foam.

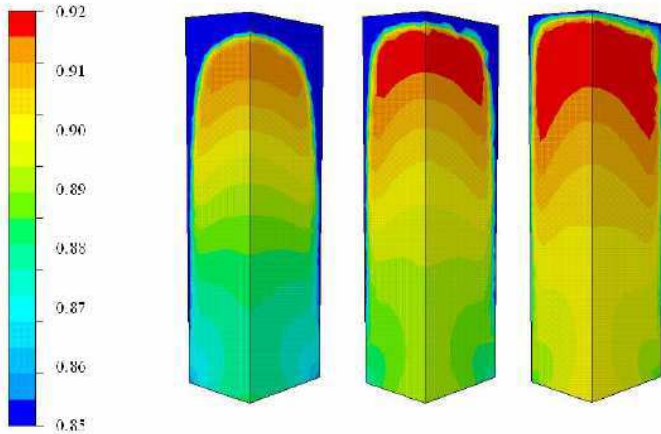


Fig 4. Results of numerical simulation with Rem3D: expansion of the PU in the closed mould after 80, 100 and 120s.

Fig. 5 shows the good agreement between the experimental and numerical velocity fields up to about 100s (after optimization of parameters of table 1).

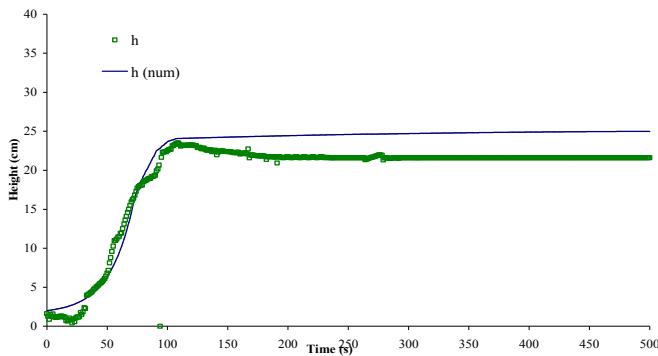


Fig. 5. Evolution of numerical and experimental height of the foam during expansion.

After 100s, agreement is less good, due to a bad description of the gas outlet during the expansion: this specific mechanism requires to improve the numerical description of the free surface of the foam, and to apply on it a more physical permeability condition. The viscoelasticity and the surface tension have also to be accounted for. Other experiments are then needed for a better estimation of the parameters.

## 4 CONCLUSIONS

In this paper, a simple experiment of expansion has been performed in order to identify the rheological parameters of the foam expansion model. The results show the good agreement between simulation and experiment up to the polymerization step, but the restrictive assumptions (no viscoelasticity, no permeability of the free surface) lead to discrepancies close to the gel point. We work now to overcome these limitations.

## ACKNOWLEDGEMENTS

The authors are grateful to Mr. P. Motte and Mr. S. Vézine from FAURECIA S.A. for their technical assistance.

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